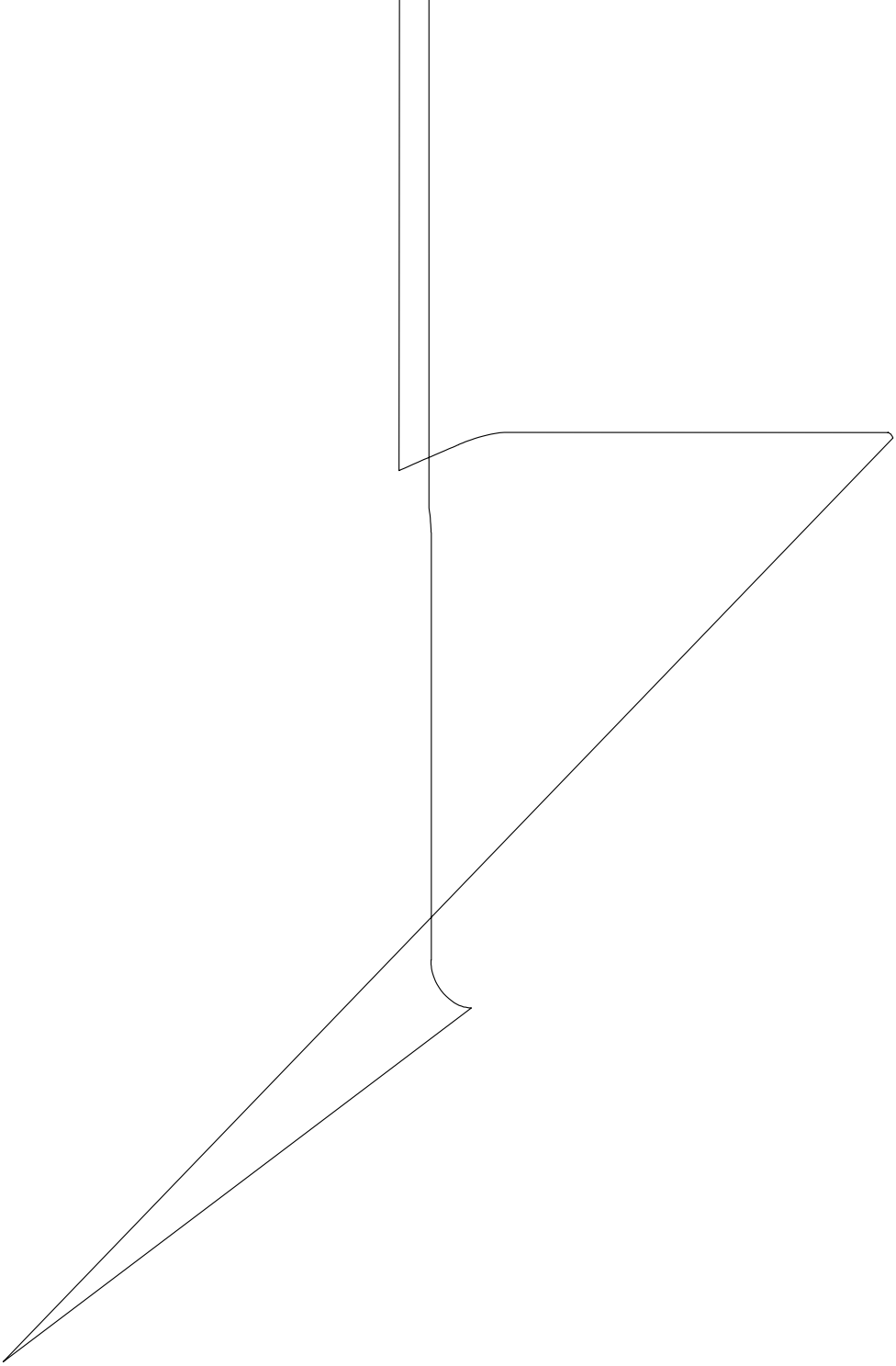

Conventions for Galvanic Cells

1. Reduction occurs at the cathode; oxidation occurs at the anode.
2. The cathode is given a positive sign (r), and the anode is given a negative

Comparison of Oxidant Strengths

- By convention, we compare redox couples on the basis of their oxidant strength; i.e., their tendency to be reduced in a redox reaction.

- Oxidant strength (or reductant strength) depends on concentration, so we define **Standard Conditions** as 1 M

Table of Oxidants

TABLE OF OXIDIZING AND REDUCING AGENTS

Oxidizing Agents (strong)	Reducing Agents (weak)

Galvanic Cell to Determine $E^{o}(Ag^{+}$

Galvanic Cell to Determine *E*